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following questions in the space provided.
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equation represent the (a) masses in
grams of all reactants and products. (b)
relative number of moles of reactants and
products. mc06se cFMs r i-vi -
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Stoichiometry SECTION 9-3 PROBLEMS
Write the answer on the line to the left.
Show all your work in the space provided.
1. 88% If the actual yield of a reaction is
22 g and the theoretical yield is 25 g,
calculate the percent yield. 2. 6.0 mol of N₂
are mixed with 12.0 mol of H₂
according to the following equation: N₂(g)
+ 3H₂(g) → 2NH₃(g) N₂; 2.0 mol a. 4798
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SECTION 2. PROBLEMS Write the answer
on the line to the left. Show all your work
in the space provided. 1. The following
equation represents a laboratory
preparation for oxygen gas: 2KClO₃(s) ...
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represent the (a) masses in grams of all
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number of moles of reactants and

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- Relationship between quantities •
Composition stoichiometry - The mass
relationships of elements in compounds
(Ch 7.3) • Reaction stoichiometry - The
mass relationships between reactants and
products in a chemical reaction Section 1
Introduction to Stoichiometry Chapter 9
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9 focuses on reaction stoichiometry: using
a balanced chemical equation to calculate

the number of grams, moles, or particles
of reactants/products involved in a
chemical reaction. Students had an
introduction to composition stoichiometry
in Chapter 3 and will now move on to
some more difficult problems. Chapter 9 -
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equation represent the (a) masses in grams
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equation represents a laboratory
preparation for oxygen gas: CHAPTER 9
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furthermore it is not directly done, Open
University Press From above we can see
that if we have 12.4 mol H₂ we need 4.13
mol N₂. We don't have that much N₂ so
the .892 mol of N₂ must be the limiting
reagent. We can now determine how much
ammonia will be produced using the mole
ratio in the balanced equation : CHEMISTRY

NOTES - Chapter 9 Stoichiometry Chapter 9 - Stoichiometry Review #1 - #18, #31, & #38 Answers . 38. To ensure that all magnesium is converted to MgO, I would use pure oxygen, not air, to carry out the reaction, because Mg could react with N₂ in air to form Mg₃N₂. The pure oxygen should be in excess. 5. a. 50 mol HI 6. a. 15.8

CHAPTER 9 REVIEW Stoichiometry SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. b The coefficients in a chemical equation represent the (a) masses in grams of all reactants and products. (b) relative number of moles of reactants and products.

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From above we can see that if we have 12.4 mol H₂ we need 4.13 mol N₂. We don't have that much N₂ so the .892 mol of N₂ must be the limiting reagent. We can now determine how much ammonia will be produced using the mole ratio in the balanced equation :

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SECTION 1 SHORT ANSWER Answer the following questions in the space provided.

1. ____ The coefficients in a chemical equation represent the (a) masses in

grams of all reactants and products. (b) relative number of moles of reactants and products.

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Chapter 9 focuses on reaction stoichiometry: using a balanced chemical equation to calculate the number of grams, moles, or particles of reactants/products involved in a chemical reaction. Students had an introduction to composition stoichiometry in Chapter 3 and will now move on to some more difficult problems.

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