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### ChemQuest 8 - PBworks

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With reference to quantum numbers, explain why the 4f sublevel can hold a maximum of 14 electrons. The 4f sublevel has the quantum numbers  $n = 4$  and  $l = 3$ . When  $l = 3$ ,  $m_l$  can only be +3, +2, +1, 0, -1, -2, and -3. Thus there are seven orbitals for the 4f sublevel. Each orbital can have electrons

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Table relating quantum numbers to orbital shape. The relationship between three of the four quantum numbers to the orbital shape of simple electronic configuration atoms up through radium (Ra, atomic number 88). The fourth quantum number, the spin, is a property of individual electrons within a particular orbital.

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